

Q1.

Which element is classified as a d block element?

- A Antimony
- B Molybdenum
- C Strontium
- D Uranium

(Total 1 mark)**Q2.**

Which element in Period 3 has the highest melting point?

- A Aluminium
- B Silicon
- C Sodium
- D Sulfur

(Total 1 mark)**Q3.**

Which ion has the largest radius?

- A F^-
- B Mg^{2+}
- C Na^+
- D O^{2-}

(Total 1 mark)

Q4.

Which element has a first ionisation energy lower than that of sulfur?

- A Chlorine
- B Oxygen
- C Phosphorus
- D Selenium

(Total 1 mark)**Q5.**

Which represents the correct order of increasing radius of the ions?

- A $F^- O^{2-} Li^+ Be^{2+}$
- B $Li^+ Be^{2+} O^{2-} F^-$
- C $Be^{2+} Li^+ F^- O^{2-}$
- D $O^{2-} F^- Li^+ Be^{2+}$

(Total 1 mark)**Q6.**

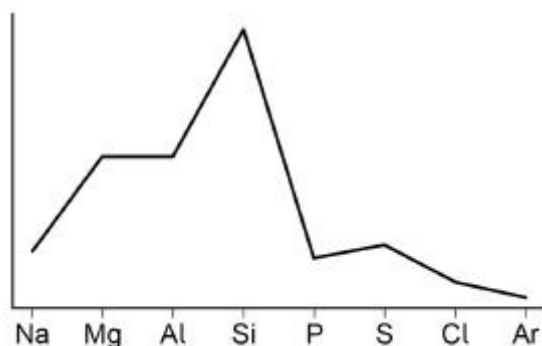
Which of these elements has the highest melting point?

- A Argon
- B Chlorine
- C Silicon
- D Sulfur

(Total 1 mark)

Q7.

The diagram shows how a property of Period 3 elements varies across the period.



What is the property?

- A Atomic radius
- B Electronegativity
- C First ionisation energy
- D Melting point

(Total 1 mark)

Q8.

Which element has the highest first ionisation energy?

- A Aluminium
- B Phosphorus
- C Silicon
- D Sulfur

(Total 1 mark)

Q9.

Which of these Period 3 elements has the highest melting point?

- A Aluminium
- B Phosphorus
- C Sodium
- D Sulfur

(Total 1 mark)**Q10.**

Which is the correct order of melting points of these Period 3 elements?

- A phosphorus > sulfur > chlorine > argon
- B argon > chlorine > phosphorus > sulfur
- C sulfur > phosphorus > chlorine > argon
- D chlorine > phosphorus > sulfur > argon

(Total 1 mark)**Q11.**

Which is the correct classification for the element yttrium (Y)?

- A s block
- B p block
- C d block
- D f block

(Total 1 mark)

Q12.

Which of the following is a correct statement about the trend in atomic radius across Period 3 of the Periodic Table?

- A** radius increases because the atoms have more electrons
- B** radius decreases because nuclear charge increases
- C** radius increases because shielding (screening) increases
- D** radius decreases because shielding (screening) decreases

(Total 1 mark)

Q13.

Which element is in the f-block of the Periodic Table?

- A** Palladium
- B** Phosphorus
- C** Platinum
- D** Plutonium

(Total 1 mark)

Q14.

Which elements are shown in increasing order of the stated property?

- A** Atomic radius: phosphorus, sulfur, chlorine.
- B** First ionisation energy: sodium, magnesium, aluminium.
- C** Electronegativity: sulfur, phosphorus, silicon.
- D** Melting point: argon, chlorine, sulfur.

(Total 1 mark)

Q15.

Which of these elements has the highest second ionisation energy?

A Na

B Mg

C Ne

D Ar

(Total 1 mark)

Q16.

Which element is in the d-block of the Periodic Table?

A Selenium

B Antimony

C Tantalum

D Lead

(Total 1 mark)